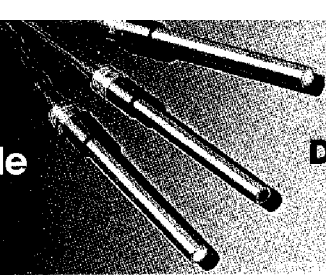


# IOTRODE™

## Ion-Selective Electrode Applications



Bulletin No. 000127

### Determination of Ionized Calcium in Skim Milk

#### INTRODUCTION

Ionized calcium in skim milk can be measured directly in the normal range (1 mM to 4 mM) utilizing the Iotron Calcium Electrode. Potassium Chloride is added to both standard and sample for consistent ionic strength.

#### EQUIPMENT

##### Meter:

pH/Millivolt meter. Readability of 1 mV required; 0.1 mV preferred. Specific ion meter will provide direct readout of final answer or concentration factor. Consult manual for millivolt measurement instructions.

##### Electrodes:

Measuring Electrode: Iotrode AB-440 Calcium ISE.  
Reference Electrodes—Double Junction—Refillable  
Calomel — RD-101 Removable Sleeve  
                  RD-201 Non-Removable Sleeve  
Ag/AgCl — RD-301 Removable Sleeve  
                  RD-401 Non-Removable Sleeve

##### Glassware:

1. 1 ml, 2 ml, and 10 ml pipets
2. 150 ml beakers
3. 100 ml graduated cylinders
4. 100 ml, 1000 ml volumetric flasks
5. Magnetic stirrer and stirbars or glass stirring rods

#### REAGENTS

- 1. 100 mM Calcium Standard Solution:** Weigh out 10.01 grams of dried (3-4 hours at 110°C) calcium carbonate into a 600 ml beaker. Add 200 ml of deionized water and stir. Continuously add 200 ml of 1 M hydrochloric acid and continue stirring until reaction is complete. Transfer quantitatively into a one liter volumetric flask and dilute to mark with deionized water. Store in a polyethylene bottle.
- 2. Ionic Strength Adjustor (ISA):** Saturate 100 ml of deionized water with potassium chloride. Reference Electrode Outer Junction Filling Solution. Dilute ISA 1:5 with deionized water.

#### ELECTRODE SET-UP

If the electrode is being used for the first time, please follow the instructions of the conditioning procedure of the electrode. Plug the sensing electrode into the G.E. or Glass jack of the meter. Fill the inner and outer chambers of the reference electrode with appropriate solutions. Plug the reference electrode into the REF or Reference jack of the meter. Technique Hints:

1. **Stirring:** Electrode response is improved if samples and standardizing solutions are stirred at a fixed rate during measurements. If magnetic stirring is not available, stir solution one minute with a clean glass stirring rod before measuring.
2. **Temperature:** The slope of the sensing electrode and the absolute potential of the reference electrode are temperature dependent. Therefore, samples and standardizing solutions should be at the same temperature.
3. **Cleaning Electrodes:** Rinse both electrodes with a fresh portion of deionized water and blot dry with tissue between all measurements.

#### ELECTRODE CALIBRATION

The primary criteria for electrode performance verification is the span test. If 25 millivolts or greater change is observed for a decade change in concentration, the performance is considered satisfactory.

1. Put 100 ml deionized water and 2 ml ISA solution into a 150 ml beaker. Place the pH meter into the MV. mode. Place electrodes in the solution to a minimum depth of one inch.
2. Pipet 1 ml of 100 mM calcium standard into the solution. Stir thoroughly. Read stable electrode potential in millivolts and record value as E1.
3. Add 10 ml of 100 mM calcium standard and stir thoroughly. Read stable electrode potential in millivolts and record as E2. Calculate S by E2 - E1. Assume S value as the slope of the electrode.

NOTE: If the slope value is below 25mV, check electrode set up and recondition sensing electrode.

**PROCEDURE**

**Sample Preparation:** Add 1 ml of ISA to 100 ml of sample.

**Sample Standardization:** Add 101 ml of prepared solution to a 150 ml beaker. Place electrodes in the solution to a minimum depth of one inch. Read stable electrode potential in millivolts and record as E3.

**Standard Addition:** Pipet 1 ml of 100 mM calcium standard into the solution. Stir thoroughly. Read stable electrode potential in millivolts and record value as E4. Calculate  $\Delta E$  by E4 - E3.

NOTE: If  $\Delta E$  is more than 15mV, **dilute standard** 1:10 with deionized water and divide final concentration by 10. If  $\Delta E$  is less than 5mV, **dilute fresh sample** 1:10 with deionized water and multiply final concentration by 10.

**CALCULATIONS**

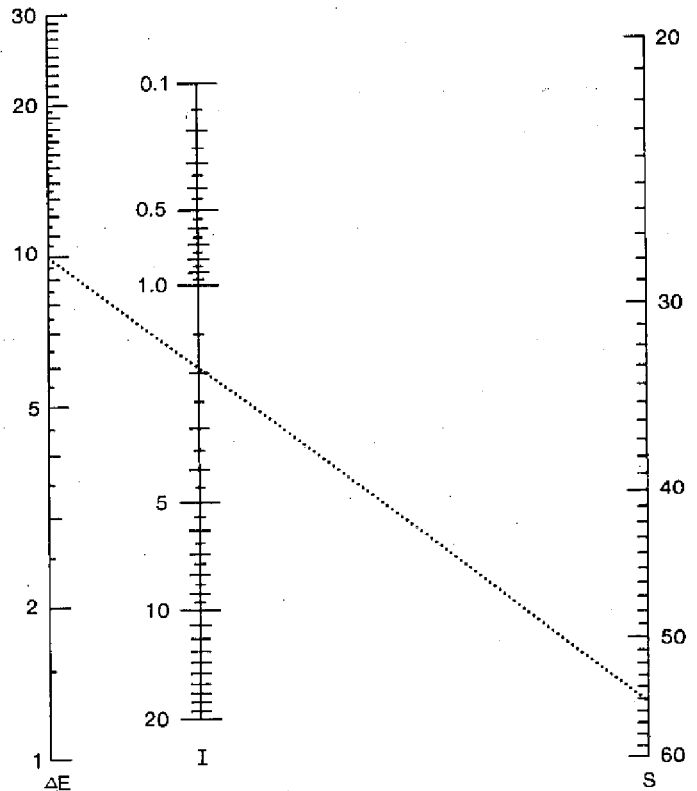
To perform the calculation for standard addition, the following factors are needed:

- $V_s$  = Sample Volume
- $V_a$  = Volume of standard added
- $C^*$  = Concentration of standard
- $\Delta E$  = Change in potential (see procedure section)
- $S$  = Slope of the electrode (see electrode calibration section)

$$\text{Concentration of Sample} = \frac{C^* \left[ \frac{V_a}{V_a + V_s} \right]}{\left[ \text{antilog} \frac{\Delta E}{S} \right] - \left[ \frac{V_s + V_a}{V_s} \right]}$$

Note: If the sample is diluted, multiply sample concentration by appropriate dilution factor.

**Standard Addition Nomograph for use with Iotrode™ Electrodes**



The standard addition nomograph method of calculation is valid when the increment of standard added to the sample is small compared to the volume. Therefore, if more than 1 ml of standard is added, please utilize the mathematical procedure above.

The following data must be known to utilize the nomograph: (See above for description of factors.)

$\Delta E$ ,  $S$ ,  $C^*$ , and  $V_s$

Utilizing  $\Delta E$  and  $S$ , draw a straight line to find  $I$  on the nomograph.

$$\text{Concentration of Sample} = \frac{(I) (C^*)}{V_s}$$

The concentration of Sample will be in the same units of concentration as  $C^*$ .

- Example:  $\Delta E = 10$  mv
- $S = 55$  mv
- $C^* = 1000$  ppm
- $V_s = 100$  ml
- $I = 2.0$  (from Nomograph)

$$\begin{aligned} \text{Sample Concentration} &= \frac{(2.0) (1000 \text{ ppm})}{100 \text{ ml}} \\ &= 20 \text{ ppm} \end{aligned}$$